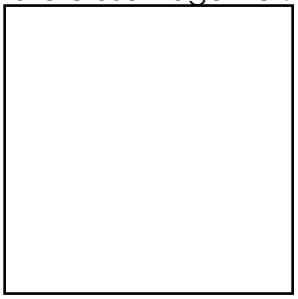
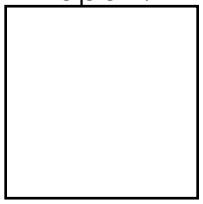
<table>
<tr><td rowspan="2">Change reactant concentration</td><td>↑ [Product] shifts eq.

to use up the extra product.</td><td>↑ [Product]</td><td rowspan="2">Change product concentration</td></tr>
<tr><td>↓ [Reactant]</td><td>↓ [Reactant] shifts eq.

to make more of the missing reactant</td></tr>
<tr><td rowspan="2"></td><td>Since heat is a product, increasing the temperature would be similar to increasing a product [ ], so eq. would shift

to use up the extra product</td><td>↑ Temperature</td><td rowspan="2"></td></tr>
<tr><td>↓ Temperature</td><td>Since heat is a reactant, decreasing the temperature would be similar to decreasing a reactant [ ], so eq. would shift

to make up for the missing reactant</td></tr>
<tr><td>Changes to pressure

(only affects gases)</td><td>If we increase the volume, the molecules are now too far apart. The equilibrium would then shift to the side of the reaction with _____ moles of a gas so the molecules are closer together.
  </td><td>↓ Pressure
(↑ Volume)</td><td>Let’s say this container contains a gas and all of the molecules are perfectly spaced apart.
  </td></tr>
</table>

Le Chatalier's Principle

How stress affects equilibrium

*Remember, a catalyst does not shift equilibrium. A catalyst simply helps a system reach equilibrium sooner.

↑ [reactant] shifts eq.

to use up the extra reactant.

Since heat is a reactant, increasing the temperature would be similar to increasing a reactant [ ], so eq. would shift

to use up the extra reactant

↓ [Product] shifts eq.

to make more of the missing product

Since heat is a product, decreasing the temperature would be similar to decreasing a product [ ], so eq. would shift

to make up for the missing product

			↑ [Reactant]
	↓ [Product]		
			↑ Temperature
	↓ Temperature		